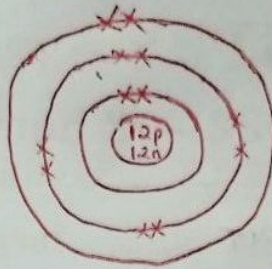


1. (a) Draw a labelled diagram showing the atomic structure of  $^{24}_{12}\text{Mg}$ . (2 marks)

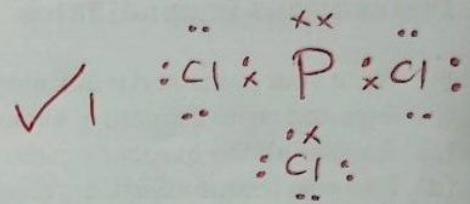
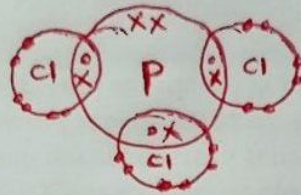
Nuclear Composition  
 $12p \checkmark \frac{1}{2}$   
 $12n \checkmark \frac{1}{2}$   
 Electrons  $\checkmark 1$



3 energy levels  $\checkmark 1$   
 Nuclear with occupied protons and neutrons  $\checkmark 1$

- (b) The atomic number of phosphorus is 15. Draw a dot (•) and cross (x) diagram for the compound formed when phosphorus reacts with chlorine, atomic number 17. (1 mark)

- 1 or 0 mark  
 - use of only dots or crosses  $\checkmark$   
 Penalties fully

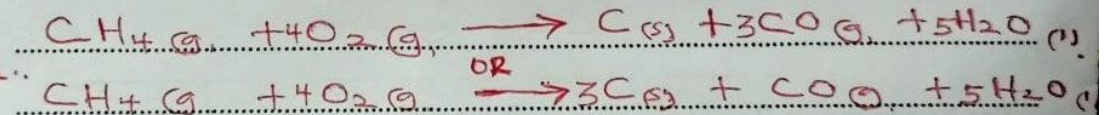


2. (a) State the condition under which a Bunsen burner produces a luminous flame. (1 mark)

When airhole/collar is closed  $\checkmark 1$  or fully closed.

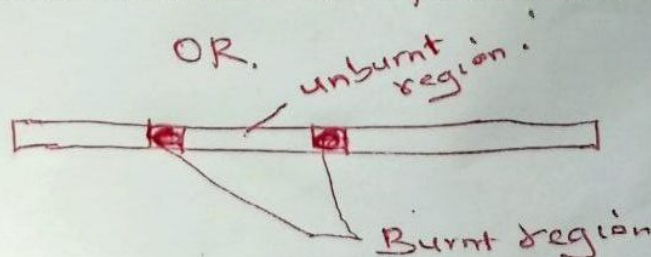
- (b) Write an equation for the reaction that takes place in a luminous flame assuming the laboratory gas is butane. (1 mark)

Wrong state symbols  
 penalty 2



- (c) One of the regions in the non-luminous flame is the unburnt gas region. Describe how the presence of this region can be shown using a wooden splint. (1 mark)

Slip a wooden splint across the middle part of the flame. The central part remains unburnt/uncharred while the outer part burns.





3. (a) The elements sodium, magnesium and aluminium belong to group I, II and III respectively. Select the element with the highest electrical conductivity and give a reason. (1 mark)

Aluminium  $\checkmark \frac{1}{2}$

It has 3 delocalised electrons while Sodium and magnesium has one and two respectively.

- (b) Complete Table 1 to show the products of electrolysis for concentrated sodium chloride and molten sodium chloride.

Table 1

Compound	Anode	Cathode
Concentrated sodium chloride	Chlorine or $\text{Cl}_2$	Hydrogen $\text{H}_2$
Molten sodium chloride	Chlorine $\text{Cl}_2$	Sodium $\text{Na}$

- Use of equation to show correct product Award

(2 marks)

4. A small piece of sodium metal was placed in a beaker containing pure water.

- (a) State two observations made during the reaction. (1 mark)

Any two correct  $\frac{1}{2}$  each.

- The piece of metal darts/floats

- melts into silvery ball.

- production of effervescence/hissing sound

- The beaker becomes warm

- (b) State and explain another observation made when a drop of phenolphthalein is added to the mixture in the beaker. (1 mark)

(marks tied)

Solution turns pink because sodium hydroxide/alkaline solution is formed.  $\checkmark \frac{1}{2}$

- (c) Explain why it is not advisable to carry out this experiment using potassium metal. (1 mark)

Potassium reacts explosively with water/more vigorously/more violent.



- Add ~~excess~~ unreacted copper
  - Filter to remove unreacted copper
  - Heat the solution to saturation
  - Allow to cool
  - Dry the crystals
5. Describe how a pure sample of copper(II) nitrate crystals can be prepared using recycled copper wire.

Concentrated nitric acid or nitric acid  
50% Nitric acid (3 marks)

- Heat the copper wire in air to form copper(II) oxide  $\checkmark \frac{1}{2}$
- Add excess copper(II) oxide to dilute nitric acid  $\checkmark \frac{1}{2}$
- Filter to remove unreacted copper(II) oxide  $\checkmark \frac{1}{2}$
- Heat the resulting solution to saturation  $\checkmark \frac{1}{2}$
- Allow it to cool to form crystals  $\checkmark \frac{1}{2}$
- Dry/filter the crystals  $\checkmark \frac{1}{2}$

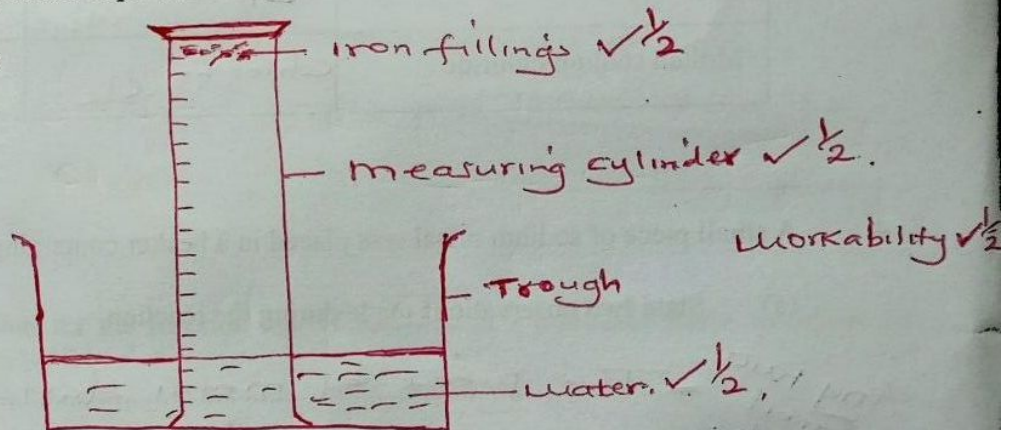
Stop marking where the candidate goes wrong.  
Use equations accepted partially.

6. The following apparatus and chemicals are used to investigate the percentage of air used when iron rusts: iron filings, 100 ml measuring cylinder, trough and water.

(a) Draw a setup of the experiment.

(2 marks)

Workability  
✓ Slanting measuring cylinder  
✓ uncalibrated in cylinder



- (b) Write an expression to show how the percentage of air used is calculated at the end of the experiment.

(1 mark)

or  
Volume instead of height

$$\frac{\text{Initial height of air column} - \text{Final height of air column}}{\text{Initial height of air column}} \times 100$$

OR

$$\frac{\text{Initial height of water} - \text{Final height of water}}{\text{Initial height of water}} \times 100$$



7. Figure 1 shows a graph of atomic radius of some group I and group II elements.

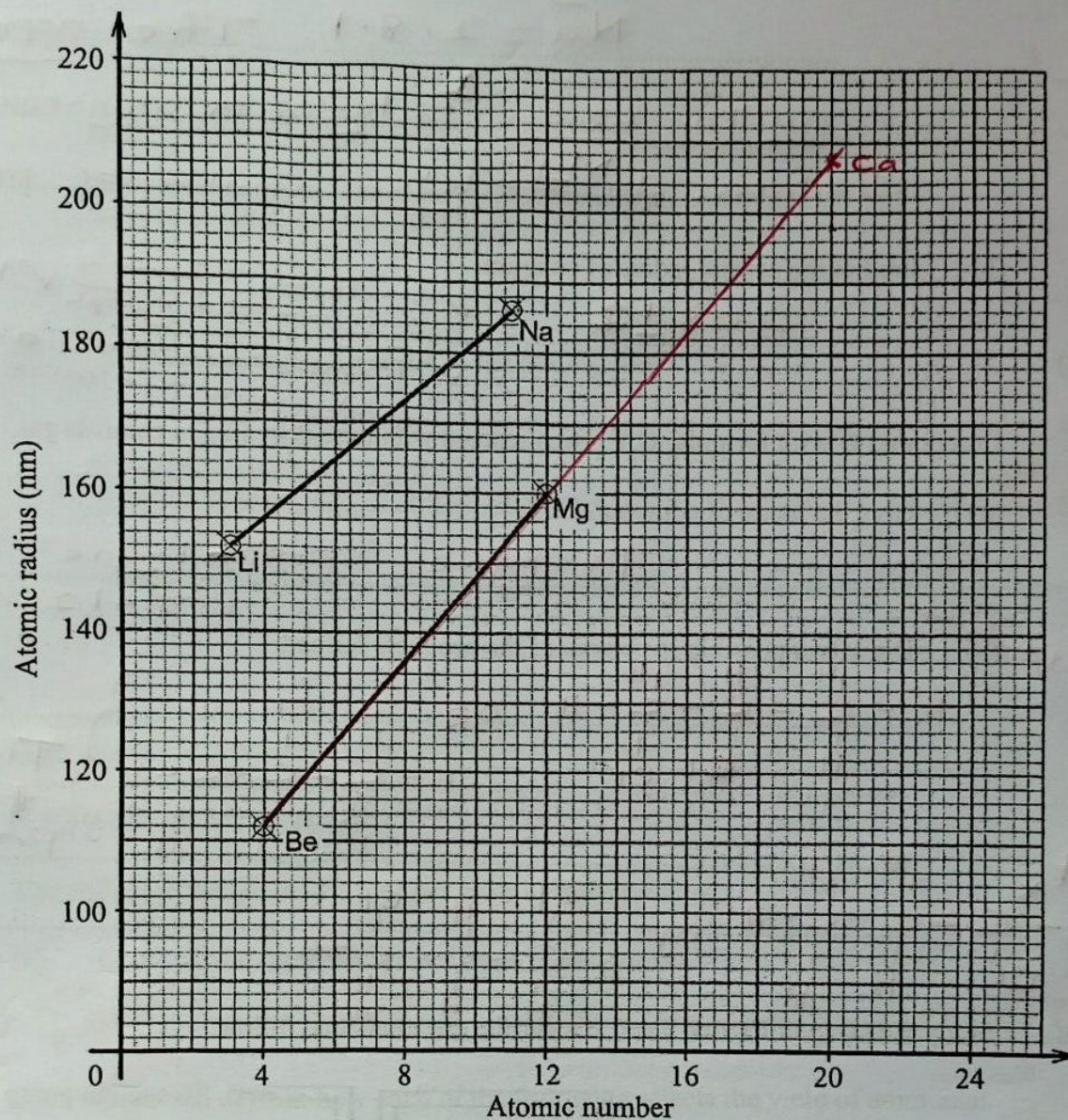


Figure 1

- (a) Explain why the atomic radius of sodium is higher than that of:

- (i) lithium.

$$Na = 2.8.1$$

$$Li = 2.1$$

(1 mark)

Sodium has 3 energy levels while Lithium has two ✓ 1 or  $Li = 2.1$  ✓  $1/2$   
 $Na = 2.8.1$



(ii) magnesium.

$$Mg = 2.8.2$$

$$Na = 2.8.1$$

The effective

nuclear charge is higher in magnesium than

Sodium. ~~It has~~ Mg has higher number of protons.

(1 mark)

(b) Predict the atomic radius of calcium.

$$208 \pm 2$$

✓ without showing on the graph ✓

✓ Extrapolate to 20 on x-axis and mark the value ✓ 1/2.

8. Compound D with formula,  $C_3H_4$ , was reacted with excess hydrogen chloride gas.

(1 mark)

(a) Give the name of compound D.

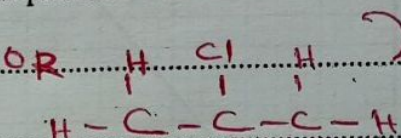
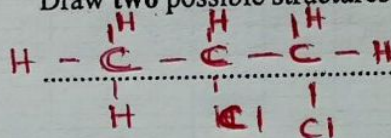
Propyne

prop-1-yne

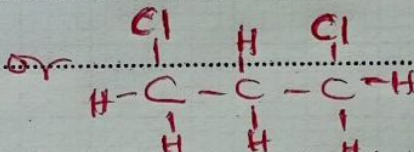
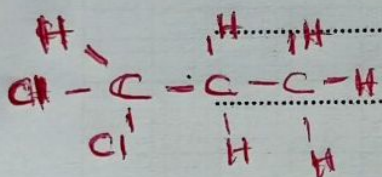
Prop-1,2-diene

(b) Draw two possible structures of the products formed.

(2 marks)



Tied to Propyne.



9. Study the setup in Figure 2 and answer the questions that follow.

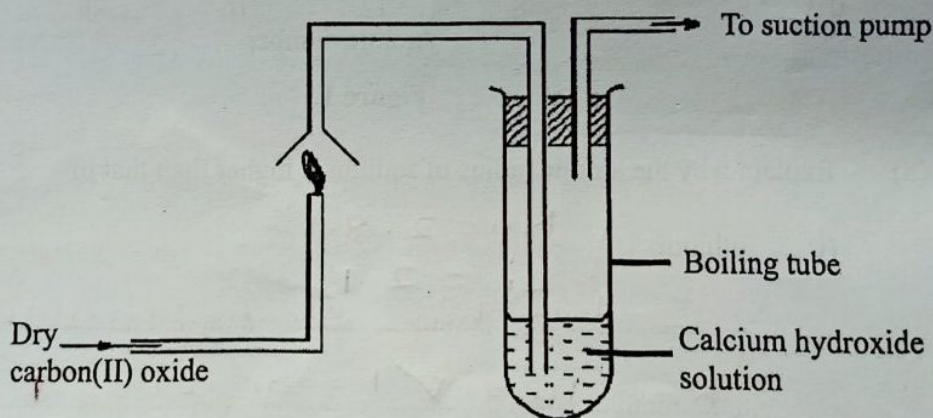


Figure 2



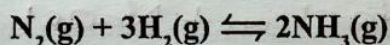
- (a) State the precaution that should be taken in carrying out the experiment. Give a reason. (1 mark)

The experiment should be carried out in a fume chamber <sup>✓ 1/2</sup> out in open Since Carbon (II) oxide is poisonous ✓ 1/2

- (b) State the observations made in the boiling tube. (2 marks)

A white precipitate is formed which dissolves ✓ 1 to form a colourless solution ✓ 1

10. Consider the following reaction:



The enthalpy change is  $-92.4 \text{ kJ}$  per mole of nitrogen.

- (a) Give the enthalpy change per mole of ammonia. (1 mark)

(missing -ve sign penalise 1/2)  $\frac{-92.4}{2} \Rightarrow -46.2 \text{ kJ mol}^{-1}$  ✓ ①

- (b) State and explain how each of the following affects the yield of ammonia:

- (i) Increase in temperature. (1 mark)

It lowers the yield of ammonia since the forward reaction is exothermic ✓ 1/2

- (ii) Finely divided iron. or backward reaction is endothermic. (1 mark)

No effect ✓ 1/2 A catalyst has no effect on the position of the equilibrium ✓ 1/2



11. Study the flow chart in Figure 3 and answer the questions that follow.

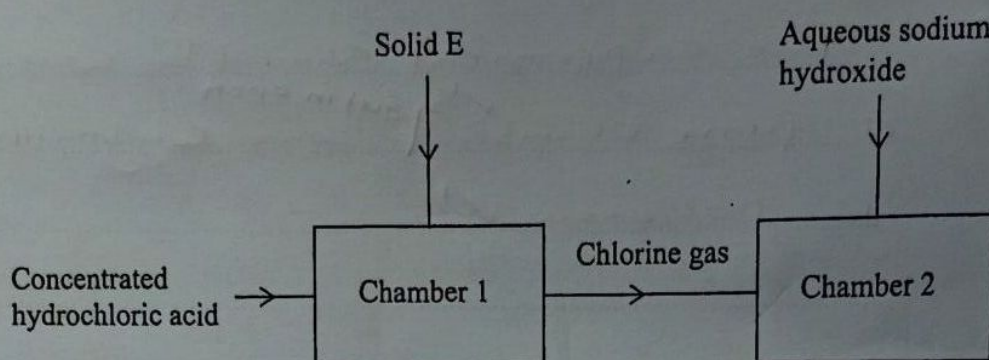
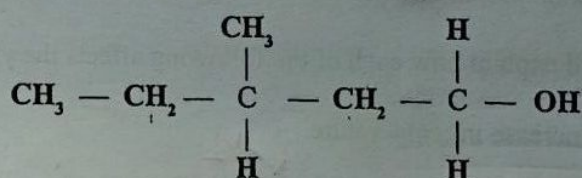


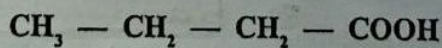
Figure 3

- (a) Identify solid E.  $KMnO_4$  /  $MnO_4^-$  or  $PbO_2$  (1 mark)  
Potassium manganate VII or Manganese (IV) oxide  
or Lead (IV) oxide
- (b) Name the type of reaction that takes place in chamber 1. (1 mark)  
Redox / oxidation
- (c) Write an equation for the reaction that takes place in chamber 2. (1 mark)  
 $Cl_2(g) + 2NaOH \rightarrow NaCl(aq) + NaOCl(aq) + H_2O(l)$

12. Compounds H and J have the following structures.



Compound H



Compound J

- (a) Give the names of:
- (i) Compound H. (1 mark)  
3-methylpentanol / 3-methylpentan-1-ol



(ii) Compound J.

(1 mark)

Butanoic acid.

(b) State the conditions necessary for H and J to react.

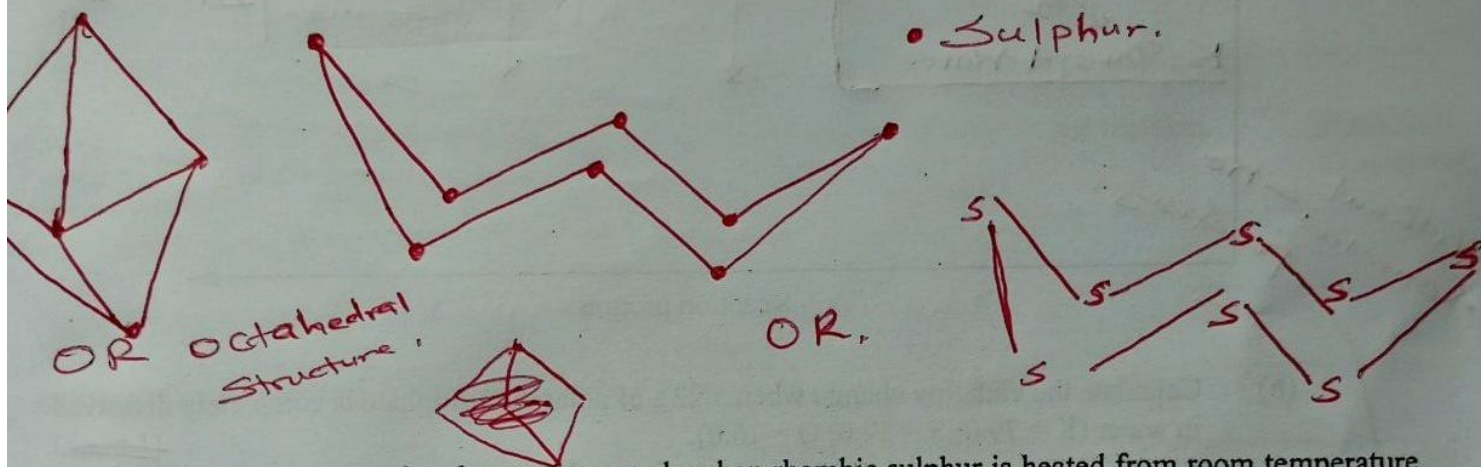
(1 mark)

- Concentrated Sulphuric VI acid ✓<sub>1/2</sub> Sulphuric VI acid  
- Warm/Heat ✓<sub>1/2</sub> Temperature between 30-60°C

13. Rhombic sulphur is one of the allotropes of sulphur.

(a) Draw the structure of rhombic sulphur.

(1 mark)



(b) Describe the observations made when rhombic sulphur is heated from room temperature until it boils.

(1 mark)

✓<sub>1/2</sub> Yellow solid forms amber liquid ✓<sub>1/2</sub>

(2 marks)

As the temp increases, the liquid becomes  
darker and viscous ✓<sub>1/2</sub>

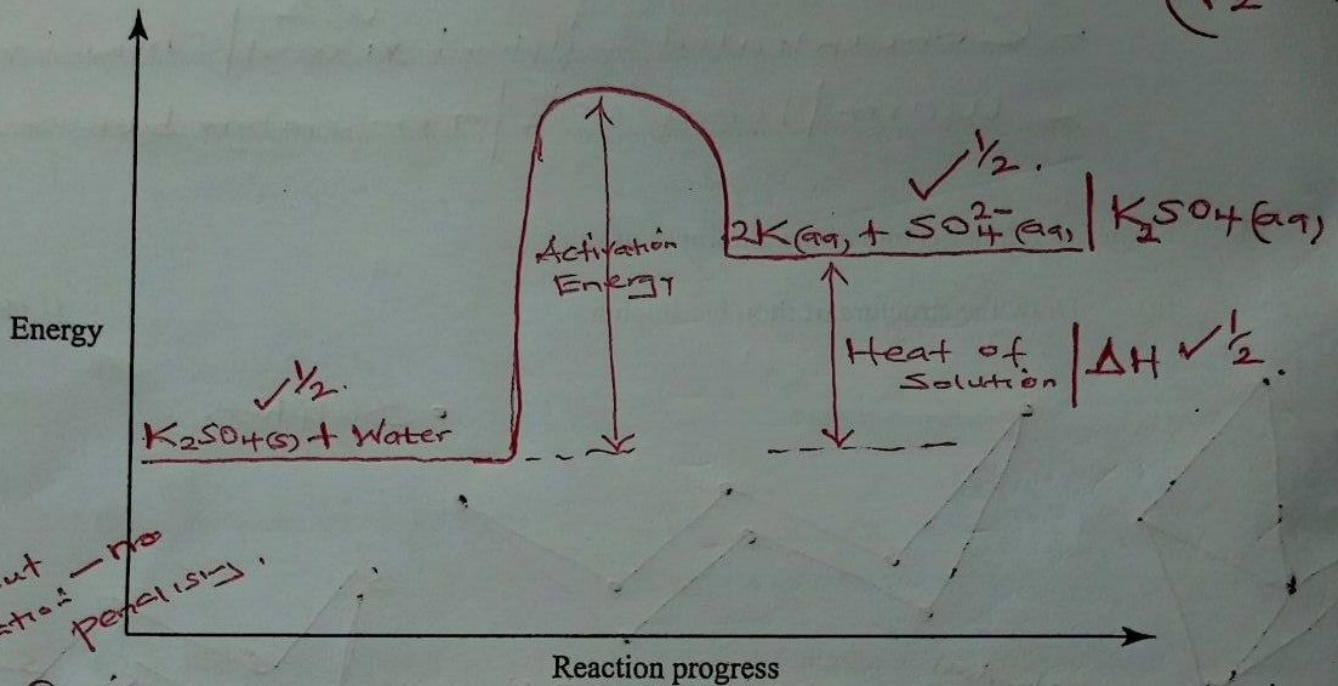
Then it turns dark red/brown and less viscous ✓<sub>1/2</sub>



14. The molar enthalpy of solution for potassium sulphate ( $K_2SO_4$ ) is +23.8 kJ.

- (a) On the axes provided, draw a labelled energy level diagram for the dissolution process of potassium sulphate in water. (2 marks)

(1 1/2 marks)



- (b) Calculate the enthalpy change when 5.22 g of potassium sulphate is completely dissolved in water (K = 39.0; S = 32.0; O = 16.0). (1 mark)

(1 1/2 marks)

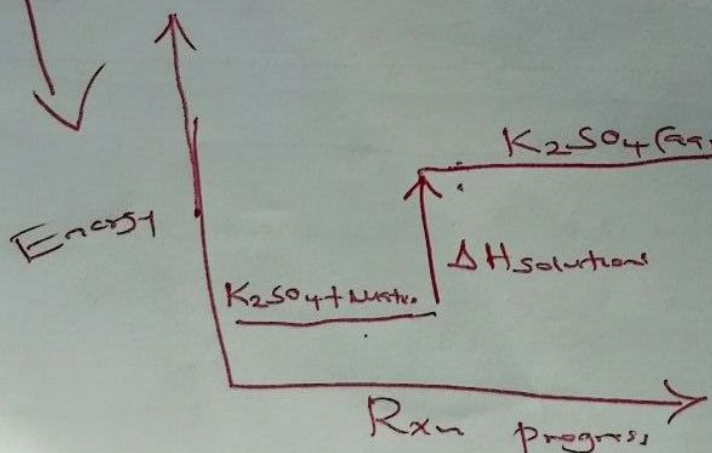
OR

$R.M.M. \text{ of } K_2SO_4 = 174$  ✓ 1/2

$\text{moles of } K_2SO_4 = \frac{5.22}{174} = 0.03$  ✓ 1/2

$\Delta H = 0.03 \times 23.8 = 0.714 \text{ kJ}$  ✓ 1/2

Wrong unit - p



For exothermic  
rxn - penalise  
fully.



15. (a) State Gay-Lussac's law.

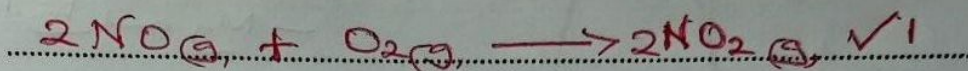
(1 mark)

When gases react, they do so in volumes that bear simple ratios to one another and to the products if gaseous at constant temperature and pressure. ✓1

- (b) 180 cm<sup>3</sup> of nitrogen(II) oxide gas was reacted with 400 cm<sup>3</sup> of oxygen gas.

- (i) Write an equation for the reaction.

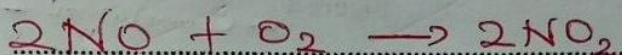
(1 mark)



- (ii) Calculate the total volume of the gases at the end of the reaction.

(3 marks)

2 marks



using ratio

$$\text{Volume of Oxygen} = 180 \times \frac{1}{2}$$

$$= 90 \text{ cm}^3 \quad \checkmark \frac{1}{2}$$

$$\text{Volume of Oxygen unreacted} = 400 - 90$$

$$= 310 \quad \checkmark \frac{1}{2}$$

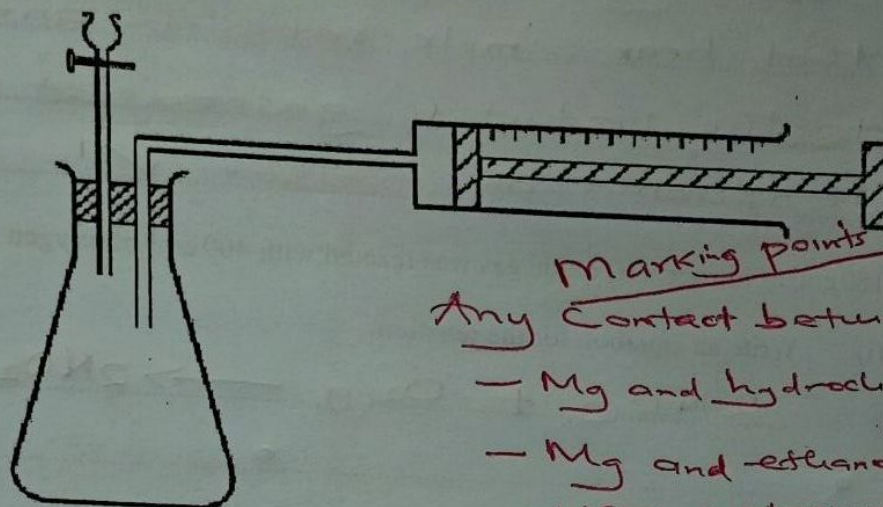
$$\text{Volume of NO}_2 = 180 \text{ cm}^3 \quad \checkmark \frac{1}{2}$$

$$\text{Total volume} = 310 + 180$$

$$= 490 \text{ cm}^3 \quad \checkmark \frac{1}{2}$$



16. Describe how the setup in Figure 4 can be used to distinguish between  $50.0\text{ cm}^3$  of  $0.2\text{ M}$  hydrochloric acid and  $50.0\text{ cm}^3$  of  $0.2\text{ M}$  ethanoic acid using pieces of  $6\text{ m}$  length of magnesium ribbon and a stop watch. (3 marks)



Marking points

Any Contact between

- Mg and hydrochloric acid ✓ $\frac{1}{2}$
- Mg and ethanoic acid ✓ $\frac{1}{2}$
- Using a stopwatch to show time difference ✓1

Figure 4

- Conclusion - HCl takes a shorter time ✓1
- Put a  $6\text{ cm}$  Mg ribbon in conical flask and add  $50\text{ cm}^3$  of HCl. Using a stopwatch, record the volume of gases collected at a time interval eg  $15\text{ cm}^3$
- Repeat the experiment using  $50\text{ cm}^3$  of ethanoic acid
- More/Higher volume of gas will be collected when HCl is used than ethanoic acid at same interval of time

OR

- The reaction ~~will~~ will take a shorter time to completion when HCl is used than when ethanoic acid is used.



17. Describe how dilute nitric(V) acid and blue litmus papers can be used to distinguish between solid samples of sodium carbonate and sodium sulphite. (3 marks)

To two different test tubes containing the samples  
 add dil. nitric(V) acid  $\checkmark \frac{1}{2}$   
 - Place wet blue litmus at the mouth of test tubes  $\checkmark \frac{1}{2}$   
 - Both turn red  $\checkmark \frac{1}{2}$   
 - Thereafter one of them is bleached  $\checkmark \frac{1}{2}$   
 - The sample that produces bleaching on the litmus is Sodium Sulphite  $\checkmark 1$

18. (a) Describe how propanone can be used to extract a pure sample of sunflower oil. (2 marks)

$\checkmark \frac{1}{2}$   
Crush the sunflower seeds using mortar and pestle  
 $\checkmark \frac{1}{2}$   
 - Add propanone and stir  
 - Decant  $\checkmark \frac{1}{2}$   
 - Leave the extract on sunlight for propanone to evaporate leaving oil behind.

- (b) State why sodium hydroxide solution is not suitable for the extraction of sunflower oil. (1 mark)

$\checkmark 1$   
It will react with oil to form soap





19.  $31.5\text{ cm}^3$  of concentrated nitric(V) acid was diluted to  $500\text{ cm}^3$ .  $10.0\text{ cm}^3$  of the dilute acid required  $25.0\text{ cm}^3$  of  $0.4\text{ M}$  sodium hydroxide for neutralisation.

(a) Calculate concentration of the:

(i) dilute acid.

(1 mark)

$$C_1V_1 = C_2V_2$$

$$0.4 \times 25 = C_2 \times 10$$

$$C_2 = \frac{0.4 \times 25}{10}$$

$$= 1\text{ M}$$

$$\text{moles of NaOH} = \frac{0.4 \times 25}{1000} = 0.01$$

$$\text{moles of HNO}_3 = 0.01$$

$$\text{Molarity of HNO}_3 = \frac{0.01 \times 1000}{10}$$

$$= 1\text{ M}$$

(ii) concentrated acid.

(1 mark)

$$C_1V_1 = C_2V_2$$

$$1 \times 500$$

$$31.5 = 15.9\text{ M}$$

(b) State the correct method for diluting the concentrated nitric(V) acid.

(1 mark)

Add acid to water ✓



20. Figure 5 shows part of a radioactive decay series.

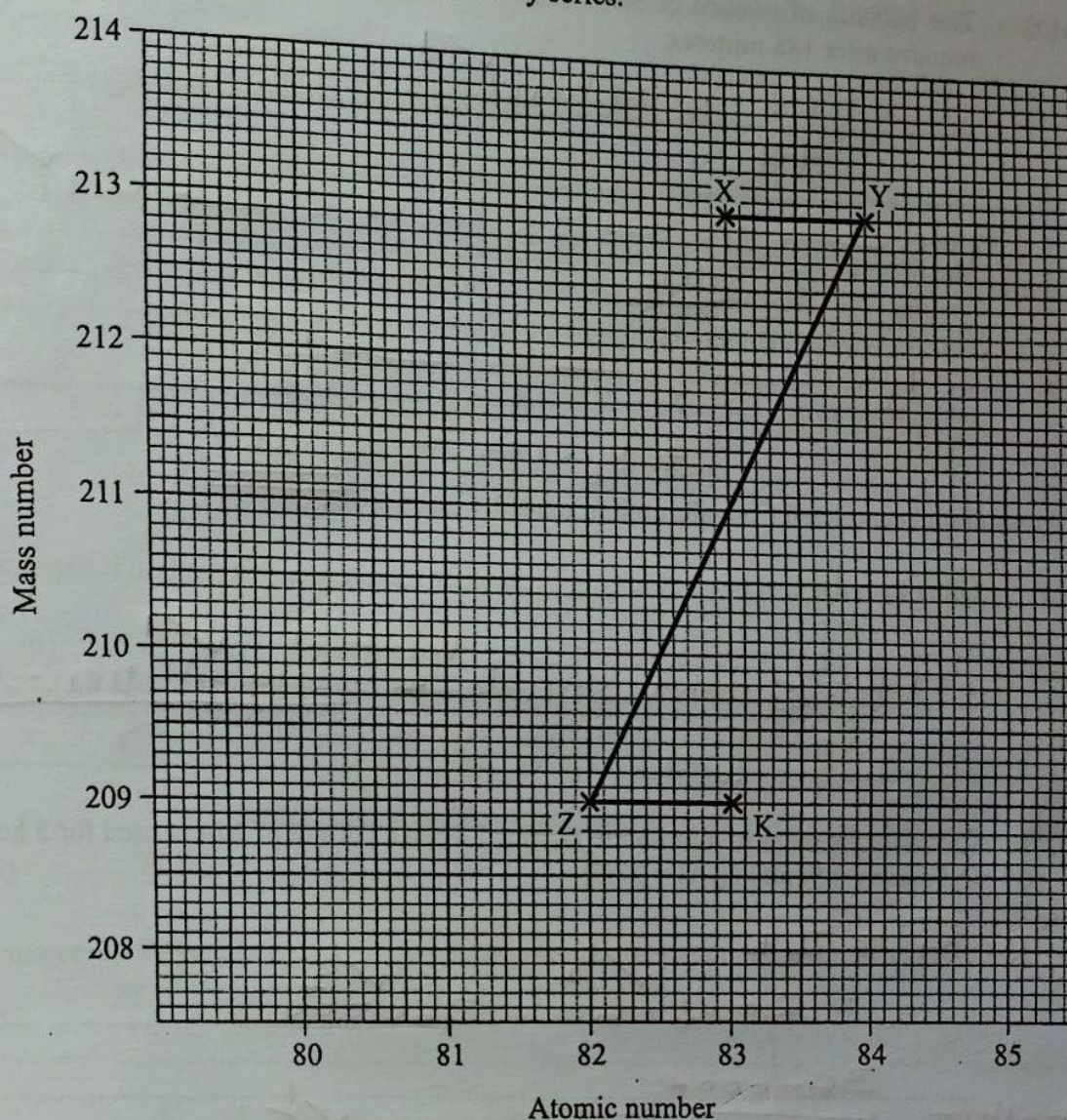
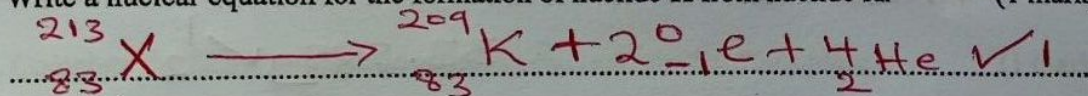


Figure 5

- (a) Write a nuclear equation for the formation of nuclide K from nuclide X. (1 mark)





- (b) The half-life of nuclide X is 47 minutes. Determine the percentage of nuclide X that remains after 188 minutes. (2 marks)

$$\frac{188}{47} = 4 \text{ half-lives} \quad \checkmark$$

$$100 \longrightarrow 50 \longrightarrow 25 \longrightarrow 12.5 \longrightarrow 6.25 \quad \checkmark$$

OR

$$X \longrightarrow \frac{1}{2}X \longrightarrow \frac{1}{4}X \longrightarrow \frac{1}{8}X \longrightarrow \frac{1}{16}X$$

$$\% = \frac{1}{16} \times 100 = \underline{\underline{6.25\%}}$$

21. Aluminium is extracted from aluminium oxide by electrolysis.

- (a) Other than the cost of electricity, give another reason why this method is expensive. (1 mark)

The graphite anode <sup>has to be</sup> replaced periodically.  $\checkmark$

- (b) Calculate the mass of aluminium obtained when a current of 20A is used for 5 hours. (1 Faraday = 96500 C; Al = 27.0) (2 marks)

$$Al^{3+} + 3e^- \rightarrow Al \quad Q = It \quad \checkmark$$

$$= 20 \times 5 \times 60 = 360000 \quad \checkmark$$

$$\text{moles} = \frac{360000}{3 \times 96500} = 1.244 \text{ mole} \quad \checkmark$$

$$\text{mass} = 1.244 \times 27$$

$$= 33.588 \text{ g} \quad \checkmark$$

22. Explain each of the following observations:

- (a) Articles made of copper turn green when left exposed in air over a long period of time. (1 mark)

Due to formation of Copper (II) Carbonate.  $\checkmark$   
 Since copper reacts with Carbon (IV) oxide /  $CO_2$ .  $\checkmark$



- (b) Addition of aqueous ammonia to a solution containing copper(II) ions produces a deep blue solution. (1 mark)

Due to formation of complex ion of tetraammine copper(II) ions ✓<sub>1/2</sub>

23. (a) State what is meant by relative atomic mass of an element.  $[\text{Cu}(\text{NH}_3)_4]^{2+}$  (1 mark)

Is the mass of one atom relative to mass of an element compared to the mass of  $\frac{1}{12}$  of Carbon-12 ✓<sub>1</sub>

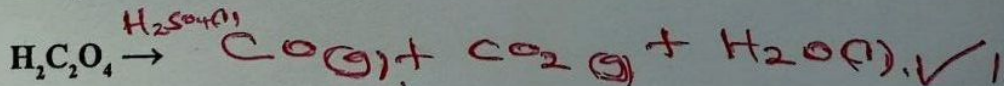
- (b) A compound of carbon and element X with formula,  $\text{CX}_4$  contains 3.6% carbon by mass. Calculate the relative atomic mass of X. (2 marks)

Let R.A.M of X be n

	<u>C</u>	<u>X</u>	
R.A.M	12	n	$\frac{96.4}{n} = 4$
% mass	3.6	96.4 ✓ <sub>1/2</sub>	$\frac{96.4}{4} = 24.1$
No. of moles	$\frac{3.6}{12}$	$\frac{96.4}{n}$ ✓ <sub>1/2</sub>	$\frac{96.4}{n} = 4 \times 24.1$
	0.3	$\frac{96.4}{n}$	$n = \frac{96.4}{24.1}$
Ratio	1	4 ✓ <sub>1/2</sub>	$\frac{96.4}{4} = 24.1$

24. Carbon(II) oxide can be prepared by dehydration of ethanedioic acid.

- (a) Complete the following equation to show the reaction that takes place. (1 mark)



State Symbols not must

- (b) Name another reagent that can be used to prepare carbon(II) oxide by dehydration. (1 mark)

- Methanoic  
- Sodium methanoate



25. Figure 6 shows an incomplete diagram of a setup for laboratory preparation of nitrogen gas.

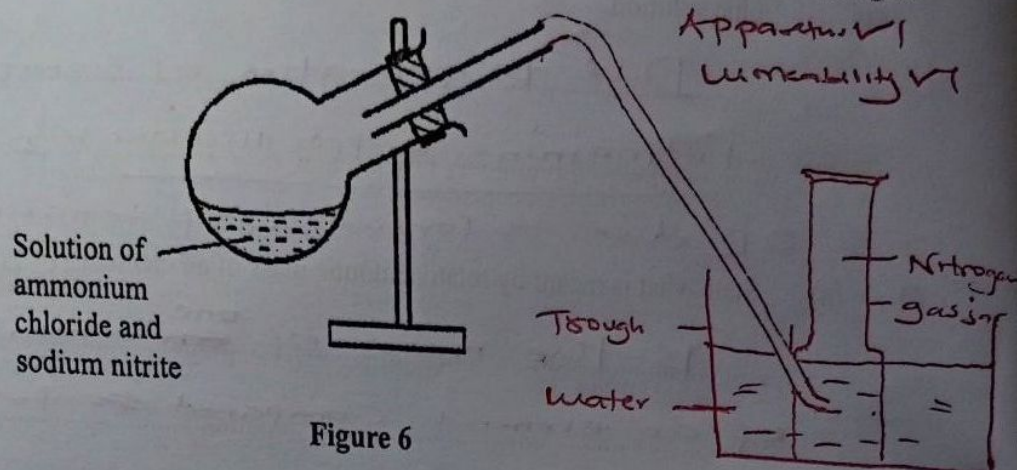


Figure 6

- (a) Complete the setup in Figure 6 to show how nitrogen gas can be collected. (2 marks)
- (b) The nitrogen prepared using this setup is purer than that obtained from air. Give a reason. (1 mark)

It has impurities such as noble gases ✓



26. Hydrazine,  $\begin{array}{c} \text{H} \quad \text{H} \\ | \quad | \\ \text{H} - \text{N} - \text{N} - \text{H} \end{array}$  is used as a fuel in rockets. Using the bond energies in Table 2, calculate the enthalpy change for combustion of hydrazine.

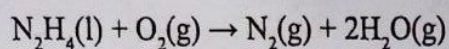


Table 2

Bond	Bond Energy kJ/mol
N—H	388
N—N	163
O=O	496
N≡N	944
O—H	463

(3 marks)

Bonds broken

$$4 \times 388 = 1552$$

$$1 \times 163 = 163$$

$$1 \times 496 = 496$$

$$\underline{2211} \checkmark$$

$$\text{Enthalpy of Combustion} = -2796 + 2211$$

$$= -585 \text{ kJ/mol} \checkmark$$

Bonds formed

$$1 \times 944 = 944$$

$$2 \times 463 = 926$$

$$\underline{-2796} \checkmark$$

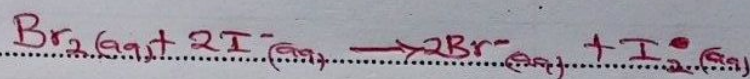


27. (a) Table 3 gives the standard reduction potentials of some group VII elements.

Table 3

Reduction equations	$E^\circ/V$
$\text{Cl}_2 + 2e \rightarrow 2\text{Cl}^-$	+1.36
$\text{Br}_2 + 2e \rightarrow 2\text{Br}^-$	+1.07
$\text{I}_2 + 2e \rightarrow 2\text{I}^-$	+0.54

State and explain the reactions that take place when aqueous bromine is added to a sample of sea water containing both chloride and iodide ions. (2 marks)



$\text{Br}_2(aq) + 2\text{Cl}^- \rightarrow$  No reaction. Bromine will oxidize iodide ions to iodine since it has more positive  $E^\circ$ . ✓

Bromine will not displace chlorine since  $E^\circ$  for  $\text{Cl}^-$  is more positive. ✓

- (b) Give a reason why potassium iodide is added to table salt.

(1 mark)

Potassium iodide is a source of iodine. Iodine is needed to regulated functioning of thyroid gland. ✓

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