



b)

	I	II	II
Final burette reading	26.0	38.30	25.30
Initial burette reading	0.70	13.00	0.00
Volume of Sol. C. used (cm ³)	25.30	25.30	25.30

Note since all the three values are the same, then they can be averaged as:

$$\frac{25.30 + 25.30 + 25.30}{3} = 25.3 \text{ cm}^3$$

i) $\frac{0.3 \times 25.3}{1000} = 0.00759$

ii) Moles of HCl in 25 cm³ of solution D = 0.00759

iii) Moles of HCl in 100 cm³ = 0.00759×4
= 0.03036